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Guidelines ... An experiment was
conducted to measure the
reaction rate of the human
salivary enzyme α -amylase. Ten
mL of a concentrated starch
solution and 1.0 mL of α -amylase

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solution were placed in a test tube. The test tube was inverted several times to mix the solution and then incubated at

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See “The Free-Response Sections
□ Student Presentation” in the AP
Physics Course Description. 4.

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show numerical results using the
value $g = 9.8 \text{ m/s}^2$, but use of 10
 m/s^2 is of course also acceptable.
Solutions usually show numerical

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Answers using both values when they are significantly different. 5.

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25 10. Mgd M M Mqu $u = + = u. 2.$
 $uq == (10^7 \sin 10^7 9.8 \text{ m s } 4.0$
 $\text{m sin}30) \text{gd} () (2) () (\square)$ For a

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Chromatography Lab. 1 point
. $u = 5.3 \text{ m s}$. Alternate solution
(kinematics method) Alternate
points . For determination of the
linear acceleration in terms of
mass and frictional force : 1 point
. $a_{FM} = 52 f$ from work shown in
part (b)

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mass and frictional force: 1 point
. $a_{FM} = 52 f$ from work shown in
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anyone help me out with the
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make sure that my answers are

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Questions Part A and B 1. Q.
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AP® CHEMISTRY 2010 SCORING
GUIDELINES Question 3 (9 points)

$8 \text{H}^+(\text{aq}) + 4 \text{Cl}^-(\text{aq}) + \text{MnO}_4^- (\text{aq}) \rightarrow 2 \text{Cl}_2 (\text{g}) + \text{Mn}^{2+} (\text{aq}) + 4 \text{H}_2\text{O} (\text{l})$
 $\text{Cl}_2 (\text{g})$ can be generated in the laboratory by

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reacting potassium permanganate with an acidified solution of sodium chloride. The net-ionic equation for the reaction is given above. (a) A 25.00 mL sample of 0.250 M NaCl reacts completely with excess KMnO_4

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Score: 10. The response earned 1 point in part (a) for correctly plotting the means on a bar graph. The response earned 1 point in part (a) for appropriate

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labels, units, and scaling. The response earned 1 point in part (a) for correctly plotting the error bars.

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Briefly explain your answer. | An unknown solid has a melting point of 112-114°C. Based on data from Table 1 above and the mixed melting point technique of this experiment, you can narrow down the possibilities to three compounds. List the compounds

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b) A mixed melting point of the unknown with acetanilide gives a melting point of 113-114°C.

Melting Point Experiment
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1. To observe the reaction to be studied, transfer 10-mL of 1.5% (0.44 M) H₂O₂ into a 50-mL glass beaker and add 1 mL of the freshly made catalase solution. The bubbles coming from the reaction mixture are O₂, which

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results from the breakdown of H₂O₂ by catalase. Be sure to keep the freshly made catalase solution on ice at all times.

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