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Calorimetry can be used to find relative specific heat values when different substances are mixed with water.

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1. Explanation: How do you think you can use calorimeters to compare the specific

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thermal abilities of substances listed on Gizmo? 2.Predict: Which substance do you think will have the highest specific heat capacity? Why? 3.Experiment: Use Gizmo to determine the final temperature for each setting listed below. Record your results in a table.

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A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to measure the heat of reaction. Acces PDF Explor ...

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When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to find a substance 's specific heat capacity . You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of

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various substances. 1.

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B 1 is a pain receptor, 2 is a temperature receptor, 3 is a light touch receptor, and 4 is a strong pressure receptor [Book]

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Correct Answer: C. The final temperature of the lead-water system will be lower than the

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final temperature of the copper-water system. A blacksmith heats a 1,540 g iron horseshoe to a temperature of 1445 ° C before dropping it into 4,280 g of water at 23.1 ° C.

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to determine the specific heat capacities of various substances. 1. On the SIMULATION pane, select Copper. Use the slider to set its Mass to 200 g. Set the Water mass to 200 g. Check that the Water temp is set to $30.0\text{ }^{\circ}\text{C}$ and the copper 's is $90\text{ }^{\circ}\text{C}$. Select the GRAPH tab, and click Play (A.

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Answers individual substances after mixing.

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Calorimetry Lab Answers Correct Answer:

A. Substance A A chemist mixes 500 g of lead at $500\text{ }^{\circ}\text{C}$ with 1,200 g of water at $20\text{ }^{\circ}\text{C}$. She then mixes 500 g of copper at $500\text{ }^{\circ}\text{C}$ with 1,200 g of water at $20\text{ }^{\circ}\text{C}$. The specific heat capacity of lead is 0.1276

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J/g ° C and the specific heat capacity of copper is 0.3845 J/g ° C.

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